

**I'm not a robot!**

Explore The first part of this section has students play around with the simulation. I do this so they can become familiar with it by making observations and have a little fun before getting into the assessment questions. As they are playing I walk around to ensure that they have turned on the symbol and abundance tabs. After about 3-5 minutes I have them start answering questions 1-6 on the first page of the lab handout. How does the number of protons change as atomic number increase by one? By one. How does the mass of the atoms change as atomic number increases by one? By one. What effect does adding a neutron have on the atom's identity? It can make the atom unstable. What effect does adding a neutron have on the atom's mass? Increase it by one. Draw the nucleus of the most abundant isotope of each of the following atoms in the boxes below. Be sure to count and label the protons and neutrons. Also show the full atomic symbol. Hydrogen has been done for you. This can be seen on the answer key I give them about 10 minutes to complete this before we go over the answers. Most students are aware of the answers to 1-4, because they have calculated atomic mass in the beginning of the year, so I do not much time on this part. Question 5-6 are the questions that students sometimes struggle with because it's a new way of representing an atom. I provide them with hydrogen and expect them to do carbon, oxygen and neon. The answers to 5 and 6 can be obtained using the simulation by observing the symbol and abundance tabs (see key). Explain As they are working on questions 1-6, I find it important to walk around to ensure they are getting the answers correct. I let them know that the element notation, showing the symbol, atomic number and mass number, will be used throughout the unit to represent isotopes and it is key they understand how to perform this task. After I've ensured that all students have completed questions 1-6, I project the answers on the board to make sure everyone has the drawing of the atoms nucleus and element notation correct. Practice After showing the correct answers, they get extra practice on writing element notation by complete the chart on the second page of the lab. To get them started I go over the first hydrogen (1st row) which can be seen on the answer key. I then ask them if they can tell me what is different about the second form of hydrogen (row 2). I am looking for students to respond with a different atomic mass, and specifically a different number of neutrons. Next I guide the conversation towards stability and ask, if Hydrogen-2 and hydrogen-3 are stable or unstable? I wait for students to respond with H-2 is stable, H-3 is unstable based on what they discovered using the simulation to answer questions 1-6. I conclude the questioning by asking, what does unstable mean? Even though most students do not understand that unstable elements are radioactive, it does begin to plant the idea in their minds that atoms can be unstable and radioactive in the process. In the event that no one answers the question correctly, I do tell them the answer. This will prepare them for later in the unit when it comes to predicting nuclear decay. As students are working on the table, I walk around and help students in need. The students that tend to need the most help are the ones that struggled in the first unit on atomic structure. This online quiz is intended to give you extra practice in writing names and notation for hundreds of elements and isotopes. Select your preferences below and click 'Start' to give it a try! The following examples illustrate entry of some common notation. Subject Scenario Type this Molecules Using subscripts to format molecular ratios in chemical formulas H<sub>2</sub>O H<sub>2</sub>O Simple Ions Entering charges Ca<sup>2+</sup> Ca<sup>2+</sup> Molecular or Compound Ions Entering charges and molecular ratios SO<sub>4</sub><sup>2-</sup> SO<sub>4</sub><sup>2-</sup> Complex Ions Grouping with subscripts and superscripts [Co(SCN)<sub>2</sub>(H<sub>2</sub>O)<sub>4</sub>]<sup>+</sup> Isotope Entering an isotopic mass number in the so-called M/A or M/Z format ^233.91Pa Chemical Reactions Entering a combination of correctly formatted chemical formulas and symbols 2 H<sub>2</sub>O<sub>2</sub> → 2 H<sub>2</sub>O + O<sub>2</sub> H<sub>2</sub>O<sub>2</sub> → 2 H<sub>2</sub>O + O<sub>2</sub> Chemical Reactions with States of Matter Entering a combination of correctly formatted chemical formulas with their respective states of matter and symbols CH<sub>4</sub>(g) + 4 S(s) → CS<sub>2</sub>(l) + 2 H<sub>2</sub>S(g) Electron Configuration Using complete notation 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>5</sup> 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>5</sup> Electron Configuration Using noble gas notation [He] 2s<sup>2</sup> 2p<sup>5</sup> [He] 2s<sup>2</sup> 2p<sup>5</sup> Equilibrium Expressions Including a stacked fraction and multiplication dots K<sub>(c)</sub> = [COCl]<sup>\*</sup>[Cl]<sup>-</sup> / [CO]<sup>\*</sup>[Cl]<sup>-</sup> Electrochemical Cell Notation Enter cell line notation. Mg(s) | Mg<sup>2+</sup>(aq) || Zn<sup>2+</sup>(aq) | Zn(s) All the topics you would expect in a worksheet that covers Isotope Notation, but this worksheet changes every time you press F9. Great for creating worksheets for independent work, or quiz make-ups or even a test. Isotope notation consists of mass number, atomic number and ion charge. Reinforce all of those plus number of protons, number of neutrons, number of electrons, cation, anion, common isotope, and atomic weight versus common mass number. Page 2 Highly visual and informative posters to engage your students in group tasks. Differentiate your student groups by giving them an option of Level 1 or Level 2 poster information. This resource contains a total of 72 element cards:- 1-30 element level 2 posters- Bromine, Krypton, Silver, Gold, Mercury and UraniumThe cards have lots of information that can be used in a multitude of ways to extend your kids in discussions, fill in the blank worksheets, tabulation tasks, graphing tasks, trend discuss The requested URL was not found on this server. Additionally, a 404 Not Found error was encountered while trying to use an ErrorDocument to handle the request. Apache/2.4.41 (Ubuntu) Server at sourcing.gftn.panda.org Port 443



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